CONTENTS

1. ELE	MENTARY	QUANTUM CHEMISTRY	1		
1.1.	PRINCI	PLES OF QUANTUM MECHANICS	1		
1.2.	APPROX	IMATE METHODS OF QUANTUM MECHANICS	10		
	1.2.1	Perturbation Theory for Stationary States	10		
	1.2.2	Perturbation Theory for Degenerate States	13		
	1.2.3	Linear Variation Method	15		
	1.2.4	Nonlinear Variation Method	17		
	1.2.5	Partitioning Method	18		
	1.2.6	Method of Effective Hamiltonian	18		
1.3.	SYMMETRY IN QUANTUM MECHANICS				
	1.3.1	Symmetry of Molecules	20		
	1.3.2	Reduction of Matrix Elements	36		
1.4.	ONE-EL	ECTRON APPROXIMATION	41		
	1.4.1	Many-Electron Systems	41		
	1.4.2	Electron Configuration	43		
	1.4.3	Matrix Elements of Operators	45		
	1.4.4	Hartree-Fock Equations	47		
	1.4.5	Density Functional Theory	49		
1.5.	FREE ATOM				
	1.5.1	Electron Configuration of a Free Atom	52		
	1.5.2	Energies of Atomic Terms	54		
	1.5.3	Spin-Orbit Interaction in Atoms	62		
1.6	CRYSTAL FIELD THEORY 63				
	1.6.1	Splitting of d-Levels in the Crystal Field	64		
	1.6.2	Weak Crystal Field	69		
	1.6.3	Strong Crystal Field	73		
	1.6.4	Improvements and Limits of the Crystal Field			
		Theory	74		
1.7	EMPIRI	CAL MODELS	77		
	1.7.1				
	1.7.2	Angular Overlap Method	80		
REFERE	NCES		86		
2. THE	NATURE	OF THE CHEMICAL BOND	89		
2.1.	MOLECU	JLAR HAMILTONIAN	89		
	2.1.1	Isolated Molecule	89		
	2.1.2	Effect of External Magnetic Field	92		
	2.1.3				
	2.1.4	Relativistic Terms	95		

2.2.	BORN-O	PPENHEIMER APPROXIMATION104				
	2.2.1	Classification of Molecular Motions104				
	2.2.2	Separation of Electron and Nuclear Motion106				
	2.2.3	Adiabatic Potential109				
2.3.	MOLECU	LAR VIBRATIONS113				
2.4.	UTILIZ	ATION OF THE ADIABATIC POTENTIAL117				
	2.4.1					
	2.4.2	Thermodynamic and Kinetic Parameters124				
	2.4.3	Molecular Mechanics128				
2.5.	JAHN-TELLER EFFECT130					
	2.5.1	Vibronic Coupling130				
	2.5.2					
		Vibronic Coupling144				
	2.5.3	The state of the s				
	2.5.4	Calculation of Vibronic Constants156				
	2.5.5	Dynamic Jahn-Teller Effect159				
2.6.	BOUND	STATES OF MOLECULES161				
	2.6.1	Physical Origin of the Chemical Bond161				
	2.6.2	Specific Features of Coordination Bond166				
2.7.	SYMMET	SYMMETRY IN CHEMICAL REACTIONS172				
	2.7.1	Woodward-Hoffmann Rules173				
	2.7.2	Frontier Orbital Theory178				
		Catalysis of Symmetry Forbidden Reactions180				
	2.7.4	Topological Approach185				
REFER	ENCES	186				
3. MO	LECULAR	ORBITAL THEORY191				
3.1.	LCAO I	APPROXIMATION191				
	3.1.1	Charge Density Bond-Order Matrix191				
		Roothaan Method for Closed-Shell Systems194				
	3.1.3	Unrestricted Hartree-Fock Method197				
	3.1.4	Restricted Hartree-Fock Method199				
	3.1.5	Half-Electron Method207				
3.2.	AB IN	ITIO APPROACH208				
	3.2.1	Basis sets208				
	3.2.2	Self-Consistent Field Procedure218				
3.3.	NON-E	MPIRICAL METHODS227				
	3.3.1					
	3.3.2	Pseudopotential Methods229				
	3.3.3	Xα Method235				
	3.3.4	NEMO and Fenske-Hall Method243				

3.4.	METHODS OF ZERO DIFFERENTIAL OVERLAP246			
	3.4.1 ZDO Approximation246			
	3.4.2 Principles of Semiempirical Parametrization248			
	3.4.3 CNDO Method258			
	3.4.4 INDO Method267			
	3.4.5 NDDO and MNDO Method276			
	3.4.6 Comparison of NDO Methods280			
3.5.	EXTENDED HÜCKEL METHOD284			
3.6.	INCLUSION OF RELATIVISTIC EFFECTS289			
	3.6.1 Dominant Relativistic Terms289			
	3.6.2 Spin-Orbit Interaction312			
3.7.	MOLECULAR ORBITAL PROPERTIES317			
	3.7.1 Canonical Molecular Orbitals317			
	3.7.2 Localized Molecular Orbitals323			
	3.7.3 Distribution of Electron Density328			
REFER	ENCES335			
4. EL	ECTRON CORRELATION345			
4.1.	CORRELATION ENERGY345			
4.2.	VARIATIONAL CONFIGURATION INTERACTION349			
4.3.	VALENCE-BOND METHOD356			
4.4.	MULTICONFIGURATION SCF METHOD360			
4.5.	MANY-BODY PERTURBATION THEORY362			
4.6.	COUPLED CLUSTER APPROACH371			
4.7.	PAIR CORRELATION METHODS373			
4.8.	GREEN FUNCTION TECHNIQUE376			
4.9.	PCILO AND PCILO/3 METHODS382			
4.10.	CIPSI METHOD388			
REFER	ENCES390			
APPEN	DICES			
1.	ATOMIC UNITS AND PHYSICAL CONSTANTS392			
2.	COORDINATE SYSTEMS392			
3.	A GUIDE THROUGH JUNGLE OF QUANTUM-CHEMICAL METHODS396			
	SUBJECT INDEX403			

LIST OF ABBREVIATIONS AND SYMBOLS

AO Atomic Orbital

AOM Angular Overlap Method AP Adiabatic Potential

CCA Coupled Cluster Approach
CFT Crystal Field Theory
CI Configuration Interaction

CI Configuration Interaction CNDO Complete Neglect of Differential OverlapP

DF Dirac-Fock
DS Dirac-Slater

DVM Discrete Variation Method

DZ Double Zeta

EHT Extended Hückel Theory
HAO Hybrid Atomic Orbital

GF Green Function

GVB Generalized Valence Bond

INDO Intermediate Neglect of Differential Overlap

JTE Jahn-Teller Effect

LCAO Linear Combination of Atomic Orbitals

LFT Ligand Field Theory
LHP Longuet-Higgins-Pople

LMO Localized Molecular Orbital
MBPT Many-Body Perturbation Theory

MC SCF Multiconfiguration Self-Consistent Field

MM Molecular Mechanics
MO Molecular Orbital

NDDO Neglect of Diatomic Differential Overlap

PCILO Perturbative Configuration Interaction using

Localized Orbitals
Primitive Function
Perturbation Theory

PZ Poly Zeta

PF

PT

REX Relativistic Extended Hückel

RHF Restricted Hartree-Fock
SCF Self-Consistent Field
SCP Slater-Condon Parameter
SOI Spin-Orbit Interaction
STO Slater-Type Orbital

SZ Single Zeta

UHF Unrestricted Hartree-Fock

VB Valence Bond

ZDO Zero Differential Overlap

```
complex conjugate of a
          vector a
          operator of the physical quantity A
Ā, <A>
          mean value of the physical quantity A
A
          vector potential of electromagnetic field
          matrix with elements Aii
          inverse matrix to A
          transposed matrix to A, (AT) = Aji
AT
A+
          hermitian conjugate to A, (A) ; = A;
|a>
          ket vector
<a |
          bra vector
          column matrix (vector)
{a}
{a}T
          row matrix (vector)
          antisymmetrization operator of n particles
An
          electron affinity of i-th orbital of atom A
          vector product of the vectors a and b
axb
a.b
          scalar product of the vectors a and b
          commutator of the operators A and B
[A,B]_
          anticommutator of the operators A and B
[A,B]_
          magnetic field induction
          speed of light in vacuum
C, Ca, Cb matrix of LCAO coefficients
          dipole moment vector
det(A)
          determinant of the matrix A
          elementary charge
E, (R)
          adiabatic potential
Eel
          electronic energy
F
          force constant matrix
          Fock operator
g
          electronic q-factor
          nuclear q-factor
g_N
g(1,2), g two-electron operator
          electron interaction matrix
          Fourier transform of the one-electron Green
G<sub>k1</sub>(ω)
          function
          supermatrix of electron repulsion integrals
Gpq,rs
          Planck constant
          reduced Planck constant (h = h/2n)
h(1)
          one-electron operator
hso
          operator of the spin-orbit interaction
```

i imaginary unit unit matrix with the elements δ_{ij} I angular momentum of a nucleus IN IA ionization energy of i-th orbital of atom A (ik|j1) two-electron integral of electron repulsion <ij|g|kl> two-electron integral of electron repulsion total angular momentum of electron ţ J Coulomb operator Jij Coulomb integral over spinorbitals exchange operator K₁ exchange integral over spinorbitals Kii Boltzman constant k 1 orbital angular momentum of electron MA mass of A-th nucleus mass of electron me m_o rest mass of a particle number of electrons n N number of atoms in a molecule Avogadro constant NA occupied molecular (spin) orbital occ P projection operator P, P^{α} , P^{β} charge density (bond-order) matrix linear momentum of a particle Pi spin-free n-particle density function Pn operator complementary to P Q R gas constant 8 spin angular momentum of electron Sij overlap integral temperature kinetic integral Tr(A) trace of the matrix A U matrix of a unitary transformation averaged integral of electron-nuclear attraction integral of electron-nuclear attraction unoccupied (virtual) molecular (spin) orbital Mulliken orbital electronegativity annihilation operator creation operator proton number of A-th nucleus core charge in units of e

```
spin-up function (m_a = +1/2)
a
ß
           spin-down function (m_a = -1/2)
           Bohr magneton
B
           nuclear magneton
\beta_N
           Dirac operators (matrices of dimension 4 x 4)
Yn
           averaged integral of electron repulsion
YAB
           representation of a symmetry point group
ra.
           direct product of representations
La o LP
\delta_{ij}
           Kronecker symbol
8(I)
           Dirac function
\Delta = 10Dq
           parameter of the ligand-field strength
           vacuum permittivity
0
           orbital energy
e i
Σ'
i,j
ηk(sk)
           summation omitting the terms with i = j
           spin function
           orbital exponent
ok(rk)
           atomic orbital
           wavelength
μ
           magnetic moment vector
            frequency of radiation
           wavenumber
           constant of the spin-orbit interaction
5(r;)
p(r)
            charge density
           conversion factor to SI units (\sigma_{SI} = e^2/4\pi\epsilon_0)
OSI
           vector of Pauli matrices (\sigma_x, \sigma_y, \sigma_z)
\sigma
            Pauli matrix of dimension 2 x 2
Pk(rk)
           molecular orbital
           electrostatic potential
\Phi_{\mathbf{u}}
           determinantal wave function (Slater determinant)
            of the electron configuration
øå
            wave function of the monoexcited configuration
oab
ij
           wave function of the biexcited configuration
           primitive basis set function
\chi_{\mathbf{k}}
\psi_{\mathbf{k}}(\mathbf{x}_{\mathbf{k}})
            spinorbital
           wave (state) function
            vector operator 'nabla'
```